

# Aquatic Chemistry: Chemical Equilibria And Rates In Natural Waters By Werner Stumm

By Werner Stumm

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Although  $K_w$  is defined in terms of the dissociation of water, this equilibrium constant expression is equally valid for solutions of acids and bases dissolved in water.  
<http://chemed.chem.purdue.edu/genchem/topicreview/bp/ch17/water.php>

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**SparkNotes: SAT Chemistry : Equilibrium - SparkNotes: Today's -**

Chemical equilibrium has been reached in a reaction when the rate of the forward reaction is Pure liquids do not appear in the equilibrium expression. Water

<http://www.sparknotes.com/testprep/books/sat2/chemistry/chapter8section2.rhtml>

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**Acid dissociation constant - Wikipedia, the free encyclopedia -**

the equilibrium of acid dissociation can be written symbolically as: Acid dissociation constants are also essential in aquatic chemistry and chemical

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<http://pasco.com/chemistry/experiments/online/chemical-equilibrium-with-soda-water.cfm>

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[http://www.academia.edu/658066/Aquatic\\_chemistry](http://www.academia.edu/658066/Aquatic_chemistry)

**The Equilibrium Constant - ChemWiki: The Dynamic -**

The equilibrium constant,  $K$  Notice the mathematical product of the chemical products raised to the which is the inverse of the autoionization constant of

[http://chemwiki.ucdavis.edu/Physical\\_Chemistry/Equilibria/Chemical\\_Equilibria/The\\_Equilibrium\\_Constant](http://chemwiki.ucdavis.edu/Physical_Chemistry/Equilibria/Chemical_Equilibria/The_Equilibrium_Constant)

**Chemical equilibrium - Wikipedia, the free encyclopedia -**

The concept of chemical equilibrium was developed after Berthollet For example, in the case of acetic acid dissolved in water and forming acetate and hydronium ions,

[http://en.wikipedia.org/wiki/Chemical\\_equilibrium](http://en.wikipedia.org/wiki/Chemical_equilibrium)

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